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162 CHAPTER 6: THERMOCHEMISTRY To convert the  
answer to joules, we write:  $101.3 \text{ J} \cdot 0.18 \text{ L atm} \cdot 1 \text{ L atm}^{-1} = -$   
 $x = w - 18 \text{ J}$  6.17 An expansion implies an increase in  
volume, therefore  $w$  must be  $- 325 \text{ J}$  (see the defining  
equation for pressure-volume work.)

### Chapter 6 Thermochemistry Test Answers

Chemistry Chapter 6 - Thermochemistry. energy. radiant  
energy. thermal energy. chemical energy. the capacity to do  
work. comes from the sun and is earth's primary energy  
source. the energy associated with random motion of atoms  
and molecules. energy stored within the bonds of chemical  
substances.

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A) exothermic, released B) exothermic, absorbed. AP Chemistry Practice Test, Ch. 6: Thermochemistry ... Start studying CHAPTER 6 THERMOCHEMISTRY TEST. Learn vocabulary, terms,

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AP Chemistry Practice Test, Ch. 6: Thermochemistry Name \_\_\_\_\_ MULTIPLE CHOICE. Choose the one alternative that best completes the statement or answers the question.  
1) A chemical reaction that absorbs heat from the surroundings is said to be \_\_\_\_\_ and has a \_\_\_\_\_  $\Delta H$  at constant pressure. A) endothermic, positive

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164 CHAPTER 6: THERMOCHEMISTRY Substituting into the above equation:  $E = 483.6 \times 10^3 \text{ J} - (8.314 \text{ J/mol} \cdot \text{K})(398 \text{ K})(+1 \text{ mol})$   $E = 4.80 \times 10^5 \text{ J} = 4.80 \times 10^2 \text{ kJ}$   
6.28 We initially have 6 moles of gas (3 moles of chlorine and 3 moles of hydrogen). Since our product is 6

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Name \_\_\_\_\_ MULTIPLE CHOICE. Choose the one alternative that best completes the statement or answers the question.

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CHAPTER 6 Thermochemistry © Houghton Mifflin Company. All rights reserved. 115 14. Two metals of equal mass with different heat capacities are subjected to the same amount of heat. Which undergoes the smallest change in temperature? a) The metal with the higher heat capacity. b) The metal with the lower heat capacity.

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Ch.6 - Thermochemistry Ch.6.1: The Nature of Energy

Energy: An object's capacity to perform work or produce

heat Potential Energy: Energy due to position or composition

(chemical bonds). Kinetic Energy: Energy due to the motion

of the object  $KE = \frac{1}{2}mv^2$  Law of Conservation of Energy:

Energy can neither be created nor destroyed,

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Chapter 6 - Thermochemistry - Mrs. Duffey - FHN  
Chapter 6 Thermochemistry flow is determined into or out of the surroundings. Because  $E_{univ} = 0$  by the first law of thermodynamics,  $E_{sys} = -E_{surr}$ ; what happens to the surroundings is the exact opposite of what happens to the system. chapter 6 test chemistry thermochemistry  
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